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Chapter 8: Chemical Reactions (Chem in 15 minutes or less) *Chapter 8 - Quantities in Chemical Reactions Ch. 8: Balancing Equations Day 1 Chapter 8 Review Chemistry Chapter 8 Extra Review Problems* **8.2 Types of Chemical Reactions** Chapter 8 - Chemical Reactions Chapter 8 Test Review **CHEMISTRY -- CH. 8 TEST REVIEW 8.1 Describing Chemical Reactions (1/2)** L30 - Ch 8 - Interpreting u0026 Balancing Chemical Equations chapter 8 reviewing key skills *How to Recognize and Classify Chemi-*

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CHAPTER 8 REVIEW Chemical Equations and Reactions SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg. 2. Consider the metals iron and silver, both listed in Table 3on page 286 of the text. Which one

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-One element in the reactant of a chemical compound replaces one element in the product of a Chemical Reaction -General Eqn= $A + Bx = AX + B$ -Ex= $Fe + Cu(NO_3)_2 \rightarrow Fe(NO_3)_2 + Cu$ -Note cations will replace cations while anions will replace anions

Chapter 8: Chemical Equations and

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Modern Chemistry 1 Chemical Equations and Reactions CHAPTER 8 REVIEW Chemical Equations and Reactions Teacher Notes and Answers Chapter 8 SECTION 1 SHORT ANSWER 1. a. d b. a c. b d. f e. e f. c 2. 8,4,9 3. a. 12 atoms b. 16 atoms c. 51 atoms d. 3 1024 atoms 4. $2\text{Al}(s) + 3\text{CuF}_2(aq) \rightarrow 2\text{AlF}_3(aq) + 3\text{Cu}(s)$ 5. $\text{NaCl}(aq) + \text{AgNO}_3(aq) \rightarrow \text{AgCl}(s) + \text{NaNO}_3(aq)$ 6. a.

CHAPTER 8 REVIEW Chemical Equations and Reactions

CHAPTER 8 REVIEW Class SEc7iB-n Chemical Equations and Reactions SHORT ANSWER Answer the following questions in the space provided. 1. Match the symbol

on the left with its appropriate description on the right. (aq) (a) (b) (c) (d) (f) A precipitate forms. A gas forms. A reversible reaction occurs. Heat is applied to the reactants.

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CHAPTER 8 REVIEW. Chemical Equations and Reactions. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. (1 point) 1. ____ A balanced chemical equation represents all the following . except (a) experimentally established facts. (b) the mechanism by which reactants combine to form products.

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CHAPTER 8 REVIEW. Chemical Equations and Reactions. SECTION 2. SHORT ANSWER Answer the following questions in the space provided. 1. Match the equation type on the left to its representation on the right. (4 points) ____ synthesis ____ decomposition ____ single-displacement ____ double-displacement. 2. ____ In the reaction described by the equation $2\text{Al}(s)$

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Chemical Equations and Reactions The evolution of energy as light and heat is an indication that a chemical reaction is taking place. CHAPTER 8 Fireworks. CHEMICAL EQUATIONS AND REACTIONS 261 SECTION 1 OBJECTIVES List three observations that suggest that a chemical reaction has taken place.

CHAPTER 8 Chemical Equations and Reactions

Chapter 8: Chemical Equations and Reactions Flashcards ... $2\text{Al} + 3\text{Fe}(\text{NO}_3)_2 \rightarrow 2\text{Al}(\text{NO}_3)_3 + 3\text{Fe}$. the replacement of a metal in a compound by a more reactive

metal, single replacement reaction.
identify

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This results in a correct chemical equation, or balanced formula equation, for the burning of methane in oxygen. $\text{CH}_4(\text{g})$

$+2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ This reaction is further illustrated in Figure 8-3. Additional Symbols Used in Chemical Equations Table 8-2 on page 246 summarizes the symbols commonly used in chemi-CHEMICAL EQUATIONS AND ...

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This results in a correct chemical equation, or balanced formula equation, for the burning of methane in oxygen. $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ This reaction is further illustrated in Figure 8-3. Additional Symbols Used in Chemical Equations Table 8-2 on page 246 summarizes the symbols commonly used in chemical EQUATIONS AND ...

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$2\text{Al}(\text{s}) + 3\text{CuF}_2(\text{aq}) \rightarrow 2\text{AlF}_3(\text{aq}) + 3\text{Cu}(\text{s})$ 5. $\text{NaCl}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$ 6. a.

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